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3-Formylindole-4-aminobenzoic Acid: A Potential Corrosion Inhibitor for Mild Steel and Copper in Hydrochloric Acid Media

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The corrosion inhibition efficiencies of Schiff base 3-formylindole-4-aminobenzoic acid (3FI4ABA) on mild steel (MS) and copper in 1.0 M HCl solution have been investigated using weight loss measurements, electrochemical impedance spectroscopy (EIS) and potentiodynamic polarization studies. The results show that inhibition efficiencies on both metals increase with increasing the concentration of the inhibitor. 3FI4ABA exhibited comparatively good inhibition towards mild steel in HCl medium even at low concentrations. The adsorption of inhibitor on the surfaces of the corroding metal mild steel and copper obeys the Langmuir isotherm. Polarization studies revealed that 3FI4ABA acts as a mixed-type inhibitor. Thermodynamic parameters (K_{ads} , ΔG_{ads}^0) were calculated using the Langmuir adsorption isotherm.

1. Introduction

Schiff bases are organic molecules possessing azomethine linkage (C=N) and have innumerable advantages, right from pharmaceutical applications to the corrosion inhibitions, in the various fields of science. The heteroatoms present in these molecules are of key importance and are responsible for the corrosion inhibition in acidic media on the metal surface [1–3]. Some recently reported Schiff bases showed effective inhibition for mild steel (MS), aluminum, copper, and zinc in acidic media [4–11].

The study on MS and copper corrosion has become much important particularly in industrial and academic fields. Enormous use of hydrochloric acid for acid pickling, descaling, and cleaning process of mild steel surface are the major reasons for the corrosion problems in the world. Increasing efforts have been made to study the corrosion behavior of metals and mechanism of inhibition during the past decade mainly through electrochemical investigations [8, 9]. Limited work has been reported for the corrosion inhibition of compounds derived from 3-formylindole in acid media. The present investigation was undertaken to examine the corrosion inhibition capacity and mechanism of inhibition of a novel heterocyclic Schiff base (3FI4ABA) molecules derived from 3-formylindole and 4-aminobenzoic acid in 1 M HCl solution on MS and Copper. The study was

performed using weight loss measurements, electrochemical impedance spectroscopy (EIS), and potentiodynamic polarization analysis.

2. Experimental

2.1. Inhibitor. Heterocyclic Schiff base was obtained by the condensation of equimolar mixture of 3-formylindole and 4-amino benzoic acid in ethanol. The reaction mixture was refluxed for 2 hours, evaporated and cooled in ice bath to obtain yellow coloured compound. Figure 1 shows the molecular structure of heterocyclic Schiff base 3FI4ABA. Anal.calcd for $C_{16}H_{12}N_2O_2$: C, 72.72; H, 4.55; N, 10.61; O, 12.12%. Found. C, 67.96; H, 5.03; N, 9.95; O, 12.12%; m.p. = 151°C; IR (KBr): $\nu_{C=N}$ = 1559 cm^{-1} . 1Hnmr : δ_{COOH} 12.15, $\delta_{CH=N}$ 9.94, δ_{NH} 5.59. $^{13}Cnmr$: δ_{COOH} 184.93, $\delta_{CH=N}$ 131.2. Mass: M^+ peak absent, base peak m/z: 144.

2.2. Solutions. The aggressive solutions of 1 M HCl were prepared by dilution of AR grade 37% of HCl (Merck) with deionized water. Inhibitor solutions were prepared in the range, 0.1 mM–1 mM concentrations.

2.3. Weight Loss Measurements. Mild steel specimens and copper specimens of dimension 1.5 × 2 × 0.1 cm were cut abraded with various grades of silicon carbide papers

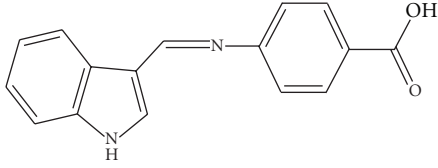


FIGURE 1: Molecular structure of 3FI4ABA.

(200, 400, 600, 800, 1000, 1200, and 2000). The exact area and thickness of each coupon were measured and washed with distilled water containing detergent. Specimens were then degreased again with acetone and finally dried. After weighing with sensitive electronic balance, specimens were immersed in 50 ml acid solutions (nondeaerated) at 28°C in the absence and presence of the inhibitor, 3FI4ABA, using hooks and fishing lines. Weight loss of metal specimens was noted for five consecutive days at 24 h interval. The experiments were carried out in duplicate, and the average values were reported. The corrosion rate (ν) is calculated by the following equation [8]:

$$\nu = \frac{W}{St}, \quad (1)$$

where W is the average weight loss of coupon, S is the total area of specimens, and t is the time of treatment (24, 48, 72, 96, and 120 hrs). The percentage of inhibition efficiency (η) is defined by [12]:

$$\eta_w\% = \frac{\nu_0 - \nu}{\nu_0} \times 100, \quad (2)$$

where ν_0 and ν are the corrosion rates of uninhibited and inhibited specimens, respectively.

2.4. Electrochemical Impedance Spectroscopy (EIS). The EIS measurements were performed in a three-electrode assembly. Saturated calomel electrode (SCE) was used as the reference electrode. Platinum electrode having 1 cm² area was taken as counter electrode. Metal specimens with an exposed area of 1 cm² were used as the working electrode. The EIS experiments were carried out on a Ivium compactSstat-e electrochemical system. 1 M HCl acid (no deaeration, no stirring) was taken as the electrolyte, and the working area of the metal specimens was exposed to the electrolyte for 1 h prior to the measurement. EIS measurements were performed at constant potential (OCP) in the frequency range of 1 KHz to 100 mHz with amplitude of 10 mV as excitation signal. The percentage of inhibitions from impedance measurements was calculated using charge transfer resistance values by the following expression [2]:

$$\eta_{EIS}\% = \frac{R_{ct} - R'_{ct}}{R_{ct}} \times 100, \quad (3)$$

where R_{ct} and R'_{ct} are the charge transfer resistances of working electrode with and without inhibitor, respectively.

2.5. Potentiodynamic Polarization. Electrochemical polarization studies of MS and copper specimens in 1 M HCl with

and without inhibitor were performed by recording anodic and cathodic potentiodynamic polarization curves. Polarization plots were obtained in the electrode potential range of -100 to +100 mV Vs corrosion potential (E_{corr}) at a scan rate of 1 mV/sec. Tafel polarization analyses were done by extrapolating anodic and cathodic curves to the potential axis to obtain corrosion current densities (I_{corr}). The percentage of inhibition efficiency ($\eta_{pol}\%$) was evaluated from the measured I_{corr} values using the following relation [13]:

$$\eta_{pol}\% = \frac{I_{corr} - I'_{corr}}{I_{corr}} \times 100, \quad (4)$$

where I_{corr} and I'_{corr} are the corrosion current densities of the exposed area of the working electrode in the absence and presence of inhibitor.

From the slope analysis of the linear polarization curves in the vicinity of corrosion potential of blank and different concentrations of the inhibitor, the values of polarization resistance (R_p) in 1 M HCl solution were obtained. From the evaluated polarization resistance, the inhibition efficiency was calculated using the relationship:

$$\eta_{Rp}\% = \frac{R'_p - R_p}{R'_p} \times 100, \quad (5)$$

where R'_p and R_p are the polarization resistance in the presence and absence of inhibitor, respectively [2].

3. Results and Discussions

3.1. Weight Loss Measurements. Weight loss of MS and copper specimens in 1 M HCl was determined at 24 h interval for a period of five days in the presence of various concentrations of inhibitor. The corrosion rates and inhibition efficiencies obtained for MS and copper specimens for a period of 24 h are listed in Tables 1(a) and 1(b), respectively. For MS specimens, the corrosion rate markedly decreased with inhibitor concentration from 0 to 0.4 mM. This response is due to the fact that the surface coverage of the inhibitor molecules on the metal through adsorption increases with the increase in concentration [14]. From 0.5 mM to 0.7 mM inhibitor concentration, percentage of inhibition efficiencies showed a gradual rise. A maximum inhibition efficiency of 88.7% was obtained for 0.7 mM. Beyond this concentration, $\eta_w\%$ got a saturation character and did not change noticeably. Figures 2 and 3 portrait the variation of rates of corrosion and inhibition efficiencies of MS specimens with immersion period. For a given period of immersion, the corrosion rate initially decreased considerably with inhibitor concentration and reached a minimum. As the concentration of 3FI4ABA increased, a marked rise in the inhibition efficiency was observed. The inhibition efficiency showed a saturation value at the inhibitor concentration 0.7 mM (critical inhibitor concentration) [15, 16].

Table 1(b) reveals the corrosion rates and inhibition efficiencies of 3FI4ABA for copper specimens in 1 M HCl. Apart from the corrosion behavior of MS specimens, copper specimens showed low corrosion rates and inhibition efficiencies for the same concentrations of inhibitor except

TABLE 1: The corrosion rate and percentage of inhibition efficiency obtained for (a) MS specimens and (b) Cu specimens immersed in 1 M HCl at 28°C for 24 h in the presence and absence of 3FI4ABA.

(a)		
C (mM)	Corrosion rate ($\text{mm}\cdot\text{y}^{-1}$)	Inhibition efficiency ($\eta_w\%$)
0	9.11	—
0.1	5.31	41.71
0.2	2.86	68.58
0.3	2.18	76.01
0.4	1.55	82.94
0.5	1.31	85.58
0.6	1.10	87.86
0.7	1.03	88.69
0.8	1.05	88.40

(b)		
C (mM)	Corrosion rate ($\text{mm}\cdot\text{y}^{-1}$)	Inhibition efficiency ($\eta_w\%$)
0	0.6002	—
0.1	0.2400	60.0
0.2	0.2315	61.4
0.3	0.2300	61.7
0.4	0.2280	62.0
0.5	0.2203	63.3
0.6	0.2187	63.6
0.7	0.2141	64.3
0.8	0.2045	65.9

for 0.1 mM. At this lower concentration, a noticeable value of $\eta_w\%$ was obtained for copper specimens, suggesting that 3FI4ABA molecules are fairly adsorbed on the surface of the metal. Beyond this concentration, only gradual rise in the inhibition efficiency was observed, and a maximum of 66% was obtained for 3FI4ABA concentration 0.8 mM. Low values of $\eta_w\%$ can be attributed to the formation of stable complex between the 3FI4ABA molecules and Cu^{2+} ions produced by corrosion. The desorption of 3FI4ABA molecules may be favored by the great affinity of inhibitor molecules to copper ions. The presence of green coloured copper complex inside the experiment vessels is the supporting evidence of this argument.

The variation of corrosion rates and inhibition efficiencies with immersion time are represented in Figures 4 and 5. Almost a similar trend to that of MS specimens was observed for copper specimens.

3.2. Comparison of Inhibition Efficiency of Schiff Base with Its Parent Amine. To compare the inhibition efficiencies of Schiff base and parent amine, weight loss measurements of MS specimens and copper specimens were performed in 1 M HCl at 28°C. The percentage of corrosion inhibition efficiencies obtained for Schiff base and parent amine on MS and copper specimens are listed in Table 2. For MS specimens, the inhibition efficiency of the Schiff base 3FI4ABA was markedly higher than that of 4-aminobenzoic acid for the studied concentrations. For copper specimens fair inhibition

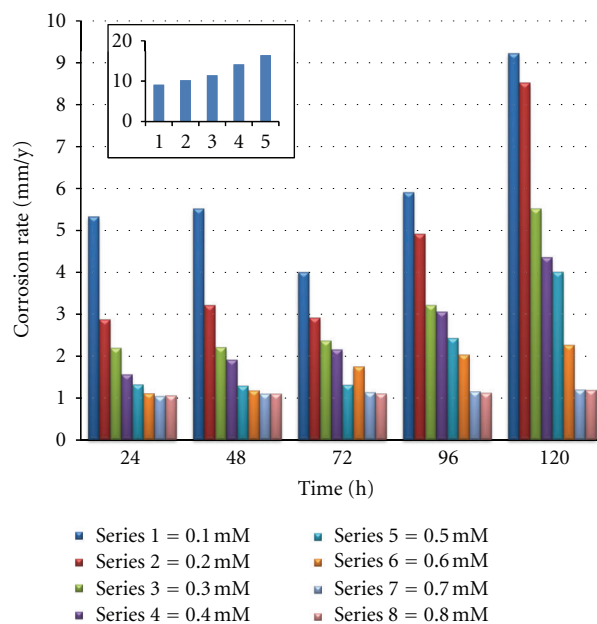


FIGURE 2: Variation of corrosion rate with immersion time of MS specimens in 1 M HCl with various inhibitor concentrations.

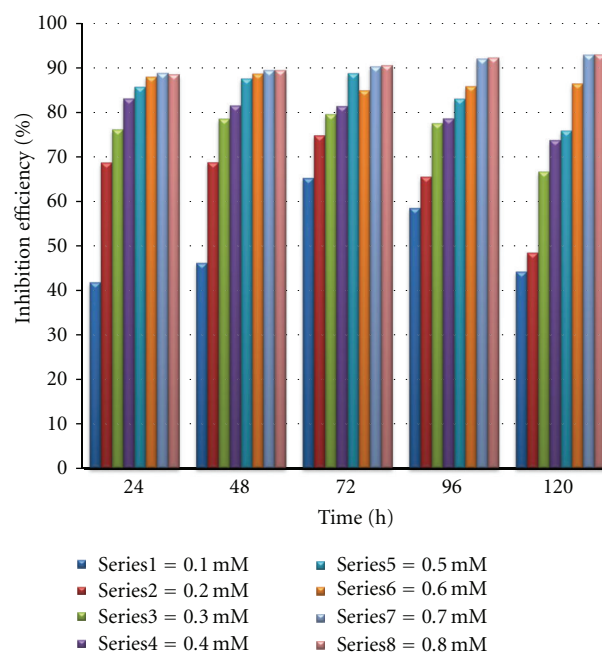


FIGURE 3: Variation of $\eta_w\%$ with immersion time of MS specimens in 1 M HCl in different inhibitor concentrations.

efficiencies were observed for the Schiff base than the parent amine. This investigation clearly establishes the role of azomethine linkage ($\text{C}=\text{N}$) present in the Schiff base, which actively participates in the corrosion inhibition mechanism.

3.3. Adsorption Isotherm and Free Energy of Adsorption. The mechanism of adsorption and the surface behavior of organic molecules can be easily viewed through adsorption

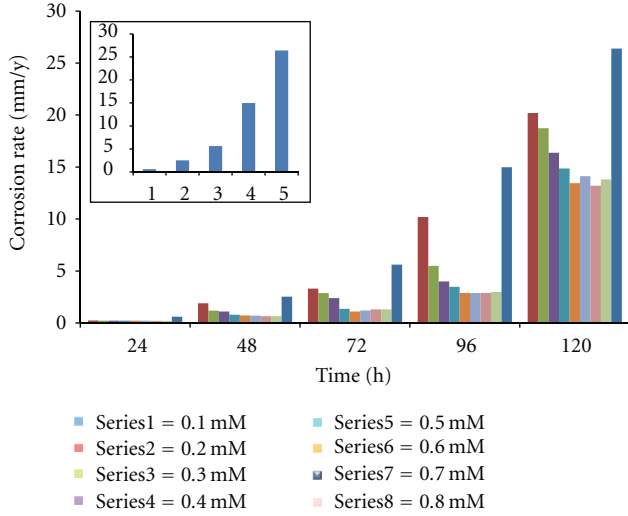


FIGURE 4: Variation of corrosion rate with immersion time of copper specimens in 1 M HCl with various inhibitor concentrations.

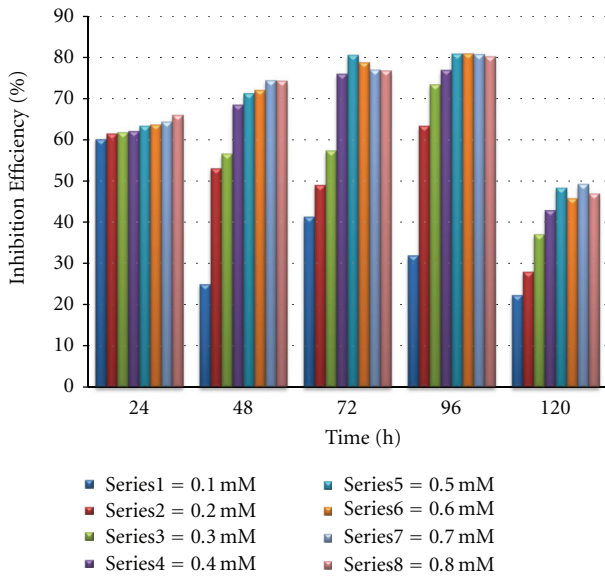


FIGURE 5: Variation of η_w % with immersion time of copper specimens in 1 M HCl with various inhibitor concentrations.

isotherms. Different models of adsorption isotherms considered are Langmuir, Temkin, Frumkin, and Freundlich isotherms. For the evaluation of thermodynamic parameters, it is necessary to determine the best-fit isotherm with the aid of correlation coefficient (R^2). Among the isotherms mentioned above, the best description of the adsorption behavior of 3FI4ABA on MS specimens and copper specimens in 1 M HCl was Langmuir's adsorption isotherm which can be expressed as

$$\frac{C}{\theta} = \frac{1}{K_{\text{ads}}} + C, \quad (6)$$

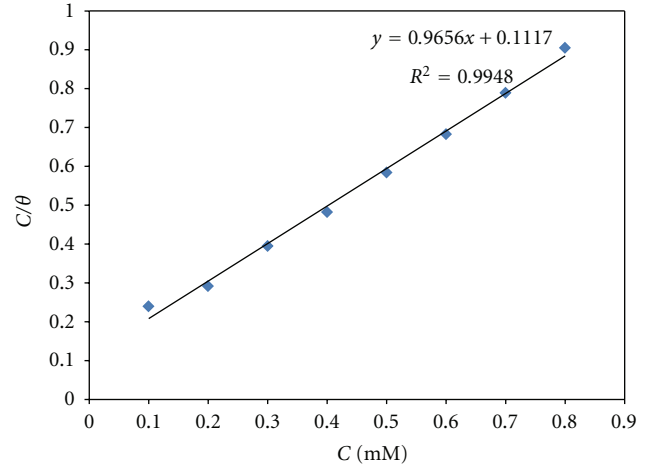


FIGURE 6: The Langmuir adsorption isotherm for adsorption of 3FI4ABA on MS surface in 1 M HCl at 28°C.

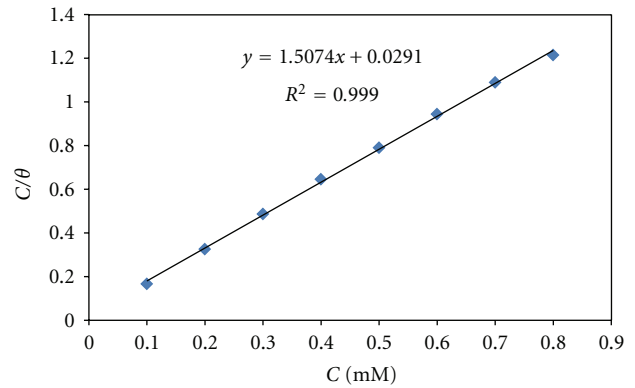


FIGURE 7: The Langmuir adsorption isotherm of 3FI4ABA on copper surface in 1 M HCl at 28°C.

where C is the concentration of the inhibitor, θ is the fractional surface coverage, and K_{ads} is the adsorption equilibrium constant [17].

Figures 6 and 7 represent the adsorption plots of 3FI4ABA obtained by the weight loss measurements of MS steel specimens and copper specimens in 1 M HCl at 28°C.

The adsorption equilibrium constant K_{ads} is related to the standard free energy of adsorption ΔG_{ads}^0 , by

$$\Delta G_{\text{ads}}^0 = -RT \ln(55.5K_{\text{ads}}), \quad (7)$$

where 55.5 is the molar concentration of water, R is the universal gas constant, and T is the temperature in Kelvin [18]. Table 3 explores thermodynamic parameters obtained from weight loss measurements of MS and copper specimens in 1 M HCl. ΔG_{ads}^0 for 3FI4ABA on MS and copper showed negative values, indicating the spontaneity of the process. The value of ΔG_{ads}^0 up to -20 kJ mol^{-1} is an indication of the electrostatic interaction of the charged molecule and the charged surface of the metal (physisorption); while ΔG_{ads}^0 is more negative than -40 kJ implies that inhibitor molecules are adsorbed strongly on the metal surface through coordinate type bond (chemisorptions [18–20]). In the present

TABLE 2: Comparison of inhibition efficiency of Schiff base and parent amine.

C (mM)	MS specimens ($\eta_w\%$)		Copper specimens ($\eta_w\%$)	
	Parent amine	Schiff base	Parent amine	Schiff base
0.2	-29.39	68.58	52.47	61.4
0.5	-18.02	85.58	54.82	65.9
0.8	-8.16	88.40	54.70	63.3

TABLE 3: Thermodynamic parameters obtained from weight loss studies.

3FI4ABA on	K_{ads}	ΔG_{ads}^0 (kJ/mol)
MS	8333	-33.2
Cu	33333	-36.3

investigation, 3FI4ABA molecules showed ΔG_{ads}^0 -33.2 and -36.3, respectively for MS and copper specimens, suggesting that the adsorption of inhibitor involves both electrostatic adsorption and chemisorption.

3.4. EIS Measurements. The corrosion response of MS and copper in 1 M HCl in the presence and absence of inhibitor has been investigated using electrochemical impedance spectroscopy at 28°C. Figures 8 and 9 represent the Nyquist plots of MS specimens and copper specimens in 1 M HCl, respectively. It is evident from the plots that the impedance response of metal specimens showed a marked difference in the presence and absence of the inhibitor 3FI4ABA.

The semicircles are observed to be depressed to the z' (real) axis or showed slight irregularity which may be attributed to the roughness or nonhomogeneous nature of the metal surface [21–24]. The capacitance loop intersects the real axis at higher and lower frequencies. At high-frequency end, the intercept corresponds to the solution resistance (R_s) and at lower-frequency end corresponds to the sum of R_s and charge transfer resistance (R_{ct}). The difference between the two values gives R_{ct} . The value of R_{ct} is a measure of electron transfer across the exposed area of the metal surface and it is inversely proportional to the rate of corrosion [25].

Impedance behaviour can be well explained by pure electric models that could verify and enable calculating numerical values corresponding to the physical and chemical properties of electrochemical system under examination [26]. The simple equivalent circuit that fits to many electrochemical system is composed of a double-layer capacitance, R_s and R_{ct} [5, 27, 28]. To reduce the effects due to surface irregularities of metal, constant phase element (CPE) is introduced into the circuit instead of a pure double-layer capacitance which gives more accurate fit as shown in Figure 10 [29].

The impedance of CPE can be expressed as

$$Z_{CPE} = \frac{1}{Y_0(j\omega)^n}, \quad (8)$$

where Y_0 is the magnitude of CPE, n is the exponent (phase shift), ω is the angular frequency, and j is the imaginary unit.

TABLE 4: Electrochemical impedance parameters of (a) MS specimens and (b) copper specimens in 1 M HCl at 28°C in the absence and presence of 3FI4ABA.

(a)				
C (mM)	R_{ct} ($\Omega \text{ cm}^2$)	C_{dl} ($\mu\text{F cm}^{-2}$)	R_s ($\Omega \text{ cm}^2$)	$\eta_{EIS}\%$
0	10.8	102	3.78	—
0.2	23	76	5.0	53
0.5	65	80.6	11.6	83.4
0.7	203.2	66.9	15.2	94.7
0.8	478	53.6	21.9	97.7

(b)				
C (mM)	R_{ct} ($\Omega \text{ cm}^2$)	C_{dl} ($\mu\text{F cm}^{-2}$)	R_s ($\Omega \text{ cm}^2$)	$\eta_{EIS}\%$
0	144	760	16.2	—
0.2	345	370	42	58
0.5	525	475	31.1	72.5
0.7	553	267	46.3	74
0.8	919	228	92.1	84

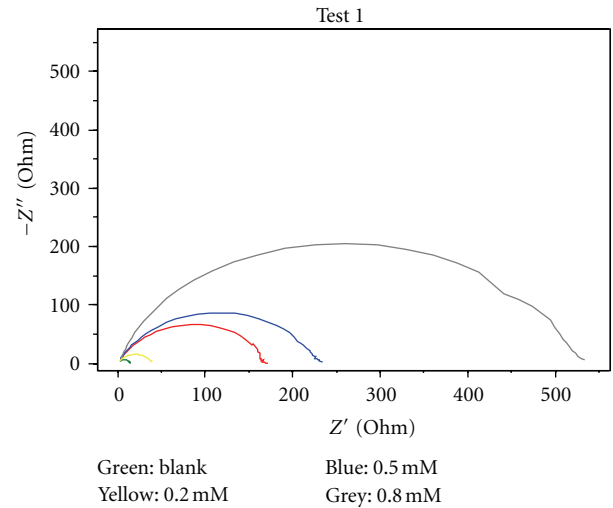


FIGURE 8: Nyquist plots for MS in 1 M HCl.

CPE may be resistance, capacitance and inductance depending upon the values of n [5]. In all experiments, the observed value of n ranges between 0.8 and 1.0, suggesting the capacitive response of CPE. The EIS parameters such as R_{ct} , R_s , and CPE and the calculated values of percentage of inhibition ($\eta_{EIS}\%$) of MS specimens and copper specimens are, respectively, listed in Tables 4(a) and 4(b).

From Tables 4(a) and 4(b) it is clear that R_{ct} values are increased with increasing inhibitor concentration. Decrease in capacitance values C_{dl} with inhibitor concentration can be attributed to the decrease in local dielectric constant and/or increase in the thickness of the electrical double layer. This emphasizes the action of inhibitor molecules by adsorption at the metal-solution interface [30]. In the case of MS specimens and copper specimens in 1 M HCl, the percentage of inhibitions ($\eta_{EIS}\%$) showed a regular increase with increasing the inhibitor concentration. A maximum of 98% inhibition

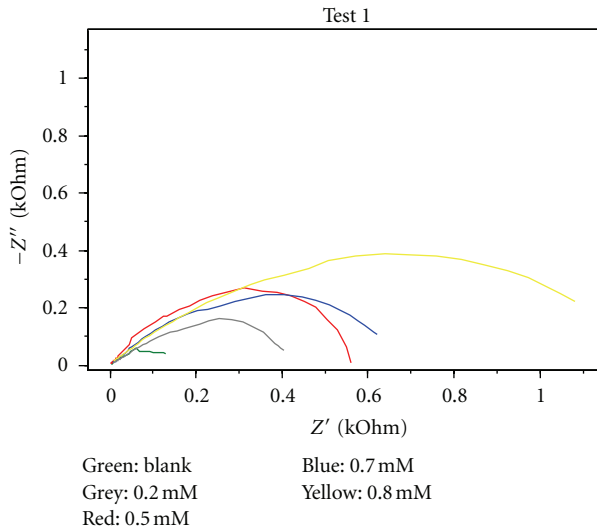


FIGURE 9: Nyquist plots for Cu in 1 M HCl.

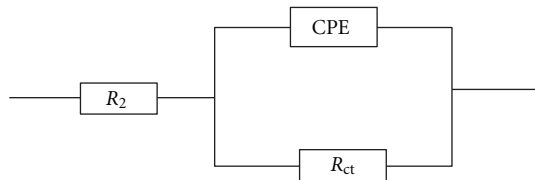


FIGURE 10: Equivalent circuit fitting for EIS measurements.

efficiency could be achieved at an inhibitor concentration of 1 mM for MS specimen. The $\eta_{\text{EIS}}\%$ data reveal that the corrosion inhibition capacity of 3FI4ABA is comparatively greater for MS than copper, which may be attributed to the high rate of desorption of the inhibitor molecules from the Cu surface.

3.5. Potentiodynamic Polarization Studies. Potentiodynamic polarization curves for 3FI4ABA in 1 M HCl for MS specimens and copper specimens are shown in Figures 11 and 12, respectively.

Polarization parameters like corrosion current densities (I_{corr}), corrosion potential (E_{corr}), cathodic Tafel slope (b_c), anodic Tafel slope (b_a), and inhibition efficiency (E_p) for MS and copper specimens are listed in Tables 5 and 6, respectively. In the case of MS specimens, a prominent decrease in the corrosion current density (I_{corr}) was observed in the presence of inhibitor 3FI4ABA. A lowest value of I_{corr} was noticed for the inhibitor solution of concentration 0.8 mM, which exhibited a maximum inhibition efficiency of 98%. Since the value of b_c changes appreciably in the presence of 3FI4ABA, it may be assumed that the inhibitor molecules are more adsorbed on cathodic sites. Generally, if the shift of E_{corr} is >85 with respect to E_{corr} of uninhibited solution, the inhibitor can be viewed as cathodic or anodic type [20, 31, 32]. In the present study, the maximum shift of E_{corr} is 14 mV, suggesting that 3FI4ABA acts as a mixed-type

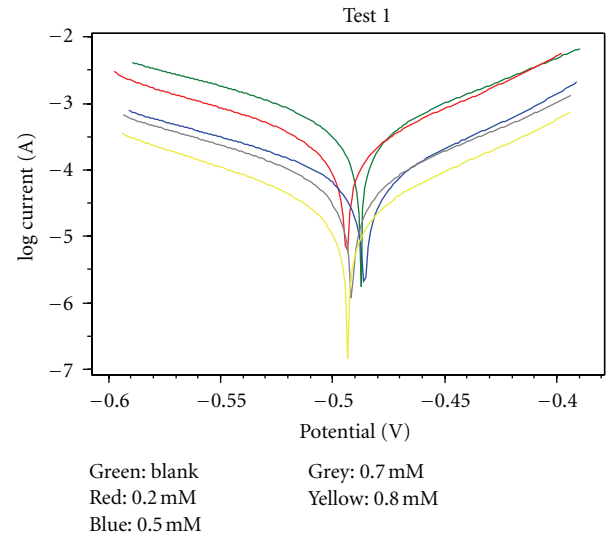


FIGURE 11: Tafel plots of MS specimens in 1 M HCl at 28°C, with and without inhibitor.

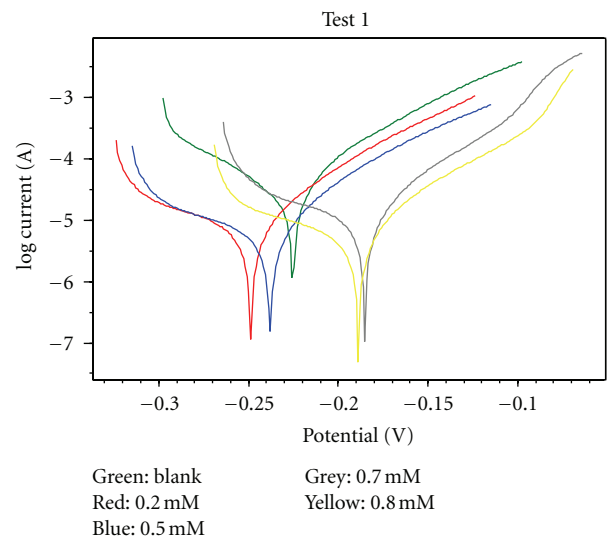


FIGURE 12: Tafel plots of copper specimens in 1 M HCl at 28°C, with and without inhibitor.

inhibitor for MS specimens in 1 M HCl. Similar results were reported by Li et al. [9].

Electrochemical polarization analysis on copper specimens showed a gradual decrease in the corrosion current densities. Inhibition efficiencies are comparable for copper specimens with the efficiencies obtained from EIS measurements. Anodic slopes (b_a) are almost remained as constant but the cathodic Tafel slopes increased prominently, suggesting that the inhibitor molecules are more adsorbed on the cathodic site.

4. Mechanism of Inhibition

MS Specimens. It is well known that the surface of the metal is positively charged in acidic media [33]. It is believed that

TABLE 5: Potentiodynamic polarization parameters of MS specimens in 1 M HCl at 28°C in the absence and presence of 3FI4ABA.

C (mM)	E_{corr} (mV/SCE)	Tafel data				Linear polarization data	
		I_{corr} ($\mu\text{A}/\text{cm}^2$)	b_c (mV/dec)	b_a (mv/dec)	$\eta_{\text{pol}}\%$	R_p (Ohm)	$\eta_{R_p}\%$
0	-478	1062	106	78	—	18	—
0.2	-492	451	114	78	57.53	42	57
0.5	-482	64	101	64	93.97	265	93.2
0.7	-489	52	95	71	95.10	337	94.7
0.8	-491	22	87	66	97.93	733	97.5

TABLE 6: Polarization parameters of copper specimens in 1 M HCl at 28°C in the absence and presence of 3FI4ABA.

C (mM)	E_{corr} (mV/SCE)	Tafel data				Linear polarization data	
		I_{corr} ($\mu\text{A}/\text{cm}^2$)	b_c (mV/dec)	b_a (mv/dec)	$\eta_{\text{pol}}\%$	R_p (Ohm)	$\eta_{R_p}\%$
0	-225	32.75	63	63	—	355	—
0.2	-191	8.76	95	50	72.3	1101	67.8
0.5	-259	7.43	99	60	77.3	1644	78.4
0.7	-252	6.57	101	63	79.8	1958	81.8
0.8	-200	5.6	111	63	82.9	2177	83.7

the Cl^- ions could be specifically adsorbed on the metal surface and creates an excess of negative charge on the surface. This will favour the adsorption of protonated Schiff base (3FI4ABAH⁺) on the surface and hence reduce the dissolution of Fe to Fe²⁺ [34]. Besides this electrostatic interaction between the protonated Schiff base and the metal surface, other possible interactions are (i) interaction of unshared electron pairs in the molecule with the metal, (ii) interaction of π -electrons with the metal, and (iii) a combination of types (i-ii) [35, 36]. If one examines the structure of Schiff base (3FI4ABA), many potential sources of inhibitor-metal interaction can be recognized. The unshared pair of electrons present on two N atoms is of key importance in making coordinate bond with the metal. The π -electron cloud of the aromatic rings and the azomethine linkage also participate in the inhibition mechanism. Furthermore, the double bonds in the inhibitor molecule permit the backdonation of metal d electrons to the π^* orbital, and this type of interaction cannot occur with amines [37]. This can be justified by the lower inhibition efficiency of the parent amine than that of Schiff base.

Copper Specimens. Anodic dissolution of copper is influenced by Cl^- concentration. At an appreciable Cl^- concentration, the dissolution of copper occurs through the formation of CuCl which will be transformed into soluble complex $[\text{CuCl}_2]^-$. The soluble $[\text{CuCl}_2]^-$ may be oxidized to Cu^{2+} ions [38]. The addition of Schiff base will cause to lower the dissolution of the metal considerably by adsorption on the surface through the potential binding surface of the Schiff base. The lower value of inhibition efficiency for copper specimens can be attributed to the formation of strong complex between the Schiff base present in the solution and the Cu^{2+} ions formed by corrosion. The higher tendency of the Schiff base molecules to combine with the Cu^{2+} ions may cause to increase the dissolution of the metal. Formation of green-coloured complex having formula $[\text{CuL}_2(\text{H}_2\text{O})_2]$

inside the corrosion vessels is the justification of this argument.

5. Conclusions

- (1) 3FI4ABA is a good inhibitor for MS in 1 M HCl. A maximum of 98% of inhibition efficiency could be achieved with this inhibitor by Tafel polarization studies.
- (2) Corrosion inhibition efficiency of the Schiff base 3FI4ABA is greater for mild steel than copper. This can be attributed to the strong tendency of Schiff base for complex formation with Cu^{2+} ions.
- (3) Compared to the parent amine, 4-aminobenzoic acid, the Schiff base 3FI4ABA exhibited higher inhibition efficiencies for MS and copper.
- (4) The inhibition mechanism is explained by adsorption. Adsorption of Schiff base on MS and copper surface obey the Langmuir isotherm.
- (5) The thermodynamic parameters of the adsorption are calculated from the adsorption isotherms which showed that both physisorption and chemisorption are involved in the inhibition process.

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